

## Chemistry Semester Two Formula Sheet

$$PV=nRT \quad R= 0.0821 \frac{L \cdot atm}{mol \cdot K}$$

$$q= c \times m \times \Delta T \quad \Delta H_{rxn} = \sum \Delta H_f^\circ \text{ products} - \sum \Delta H_f^\circ \text{ reactants}$$

$$M = \frac{\text{mol of solute}}{\text{L of Solution}} \quad M_1 V_1 = M_2 V_2$$

$$aA + bB \leftrightarrow cC + dD \quad K_{eq} = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

$$pH = -\log [H^+] \quad [H^+] = 10^{-pH} \quad pH + pOH = 14$$

$$C_B M_A V_A = C_A M_B V_B$$